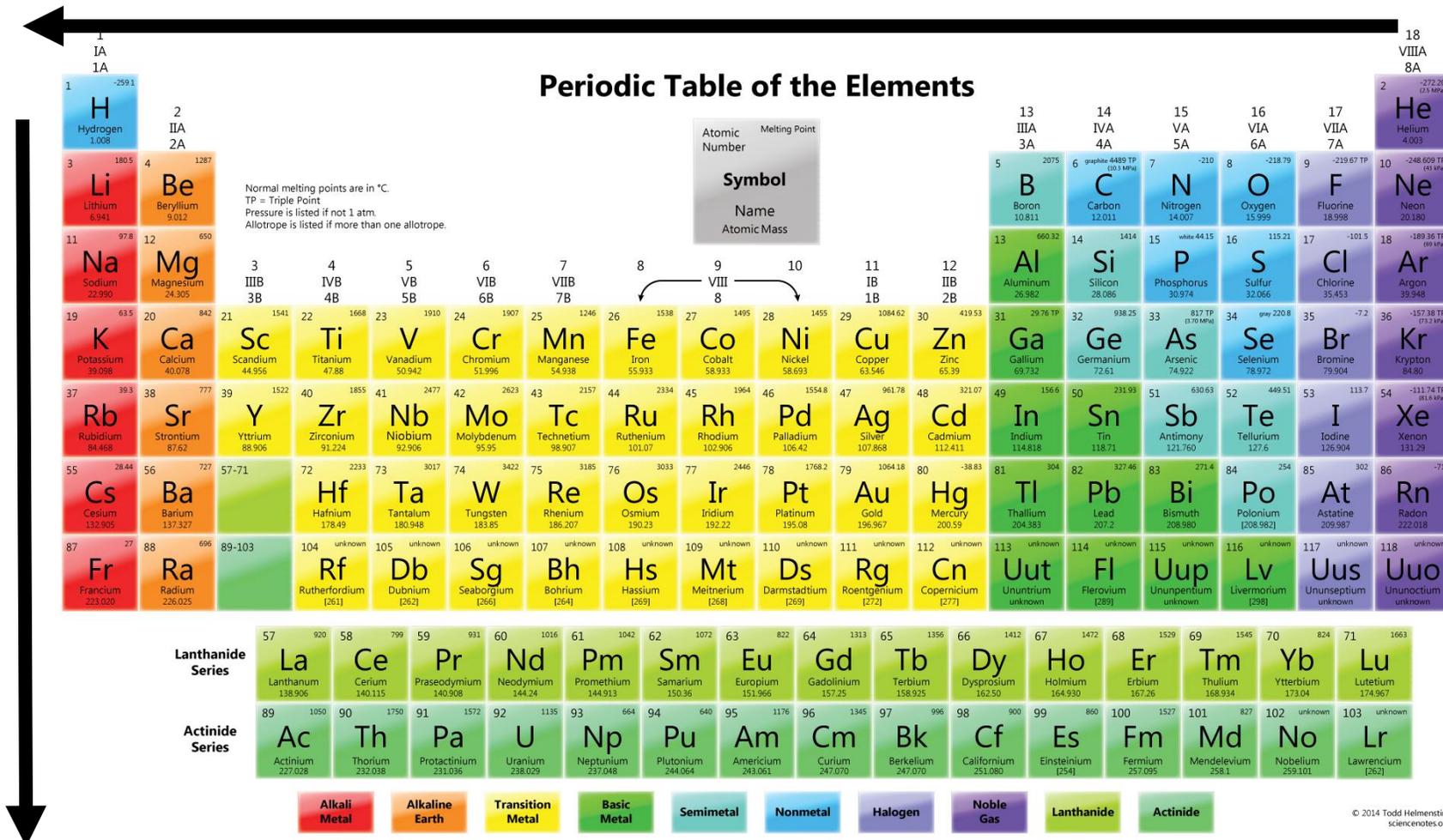


# [3.4] The Periodic Table and Periodic Trends



# Father of the Periodic Table

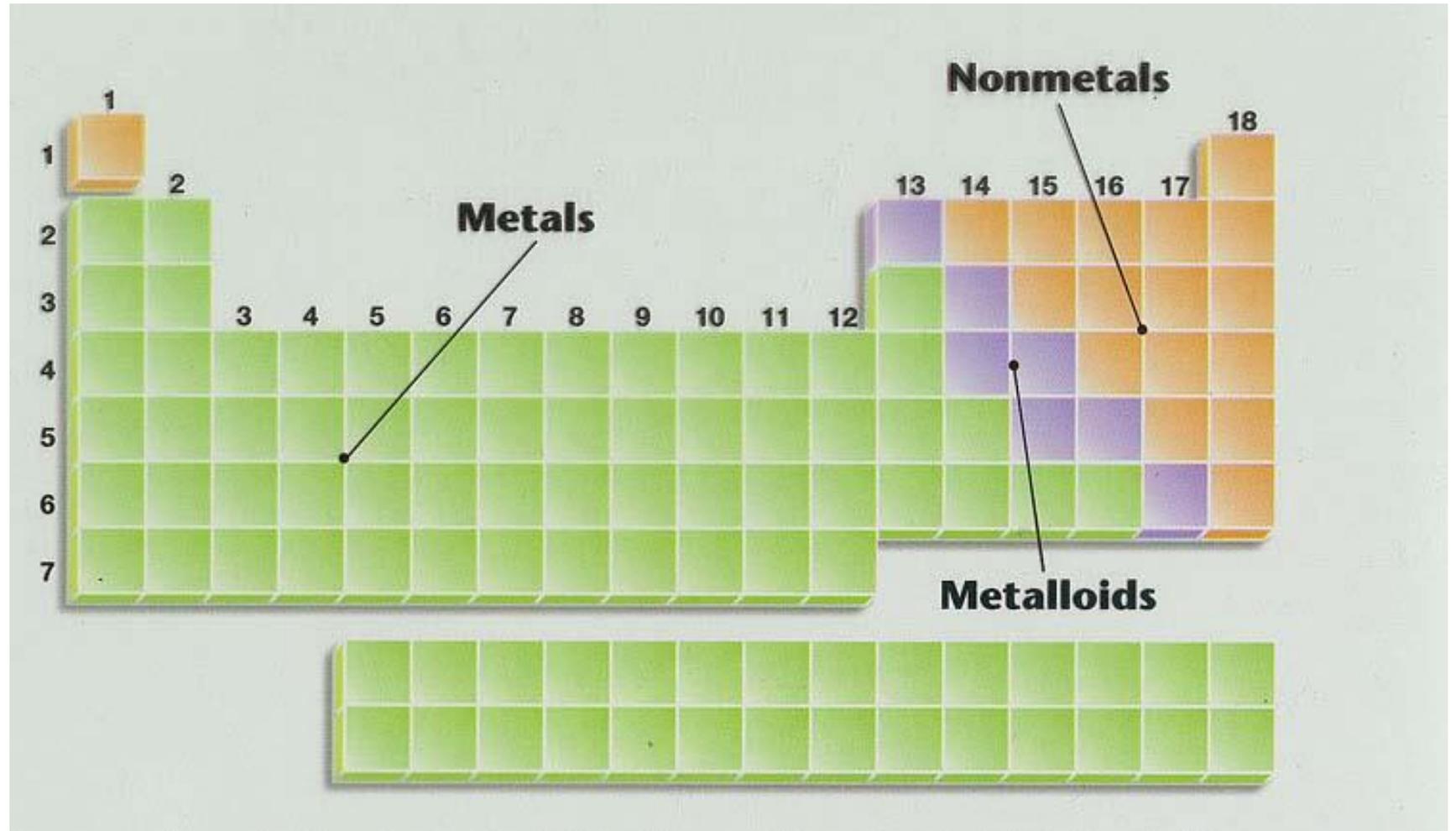
- **Dmitri Mendeleev:** Scientist who did a lot of work in the development of the modern periodic table
- Early periodic tables were arranged by **atomic masses** but Mendeleev was the first to arrange it by **atomic numbers** (number of protons) which became the **modern periodic table**

1 H 1.008																	2 He 4.003	
3 Li 6.941	4 Be 9.012											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 18.99	10 Ne 20.18	
11 Na 22.99	12 Mg 24.31											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95	
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.88	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.71	29 Cu 63.55	30 Zn 65.38	31 Ga 69.72	32 Ge 72.64	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80	
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc 98.91	44 Ru 101.07	45 Rh 101.07	46 Pd 106.36	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.76	52 Te 127.60	53 I 126.91	54 Xe 131.29	
55 Cs 132.91	56 Ba 137.33	57-70 * Lanthanide series	71 Lu 174.97	72 Hf 178.49	73 Ta 180.95	74 W 183.85	75 Re 186.21	76 Os 190.23	77 Ir 192.22	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po 209	85 At 210	86 Rn 222
87 Fr 223	88 Ra 226	89-102 ** Actinide series	103 Lr 260	104 Rf 261	105 Db 262	106 Sg 263	107 Bh 264	108 Hs 265	109 Mt 266	110 Uun 267	111 Uuu 268	112 Uub 269	114 Uuq 270					

* Lanthanide series	57 La 138.91	58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm 145	62 Sm 150.36	63 Eu 151.96	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.05
** Actinide series	89 Ac 227	90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np 237	94 Pu 244	95 Am 243	96 Cm 247	97 Bk 247	98 Cf 251	99 Es 252	100 Fm 257	101 Md 258	102 No 259

# The Periodic Table

The periodic table can be divided into **metals**, **non-metals** & **metalloids**



# Metals

- Found on the **left & middle** of the periodic table
- **Shiny, metallic lustre**
- Good **conductors** of heat and electricity.
- Usually **flexible** when in sheet form
- **Malleable & ductile**
- **Solid** at room temperature (except Hg) because of **high melting point** and **boiling point**



# Non-metals

- Found on the **right side** of the periodic table
- Solids are **dull** to lustrous
- **Poor conductors** of heat and electricity
- **Gas, liquid or brittle solids** at room temperature

**Carbon**



**Phosphorus**



**Sulfur**



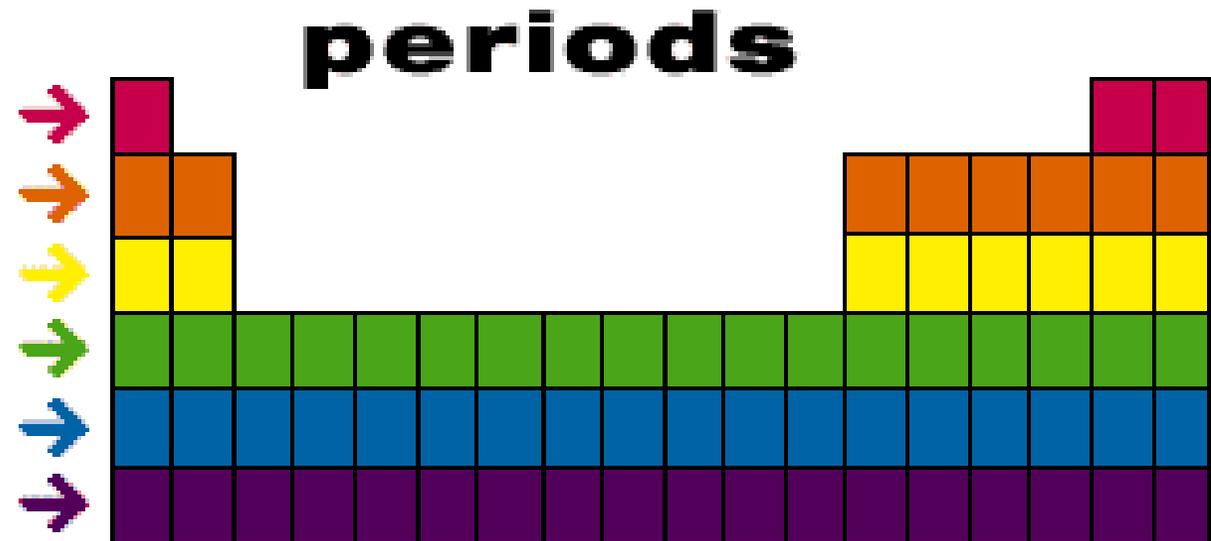
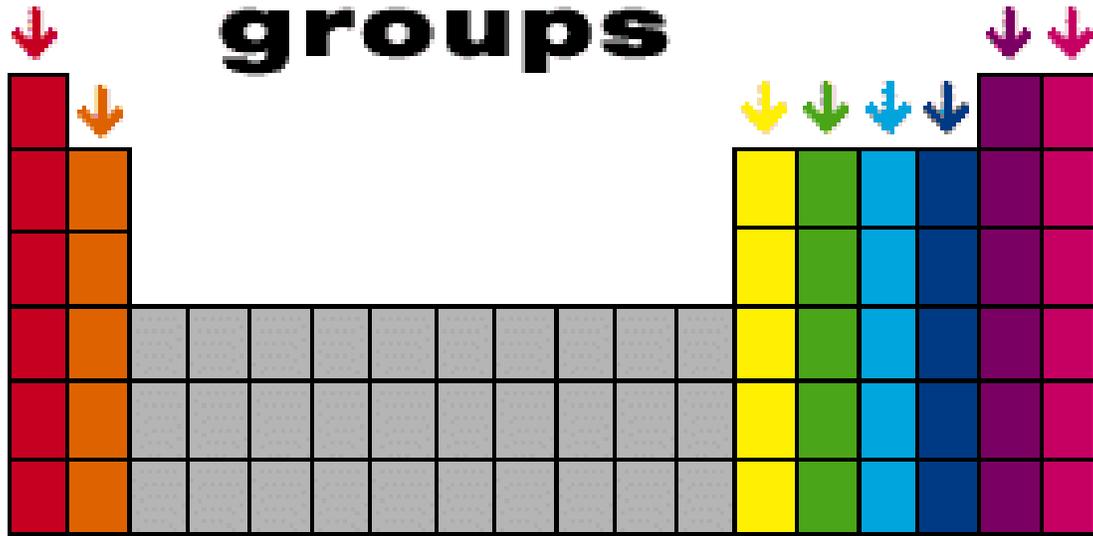
# Metalloids

- Includes **B, Si, Ge, As, Sb, Te, Po**
- They are the elements found along the **stair-step line** that separates the metals & non-metals
- Metalloids are also known as "*semiconductors*"
- Metalloids have properties of **both metals & non-metals**



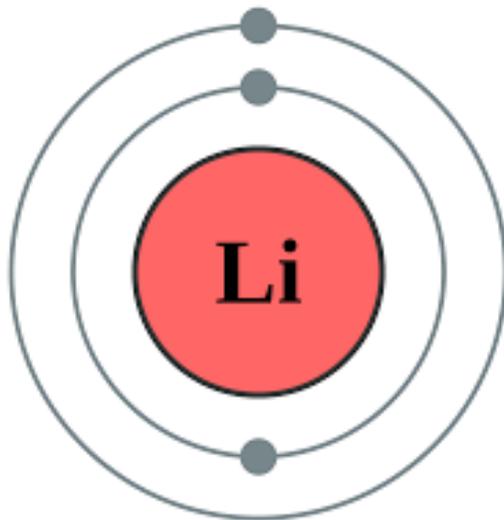
# Introduction to the Periodic Table

- The periodic table is organized into **vertical columns (group or family)** and **horizontal rows (period)**
- Each **group/family** share **similar chemical and physical properties**



# Group 1: Alkali Metals

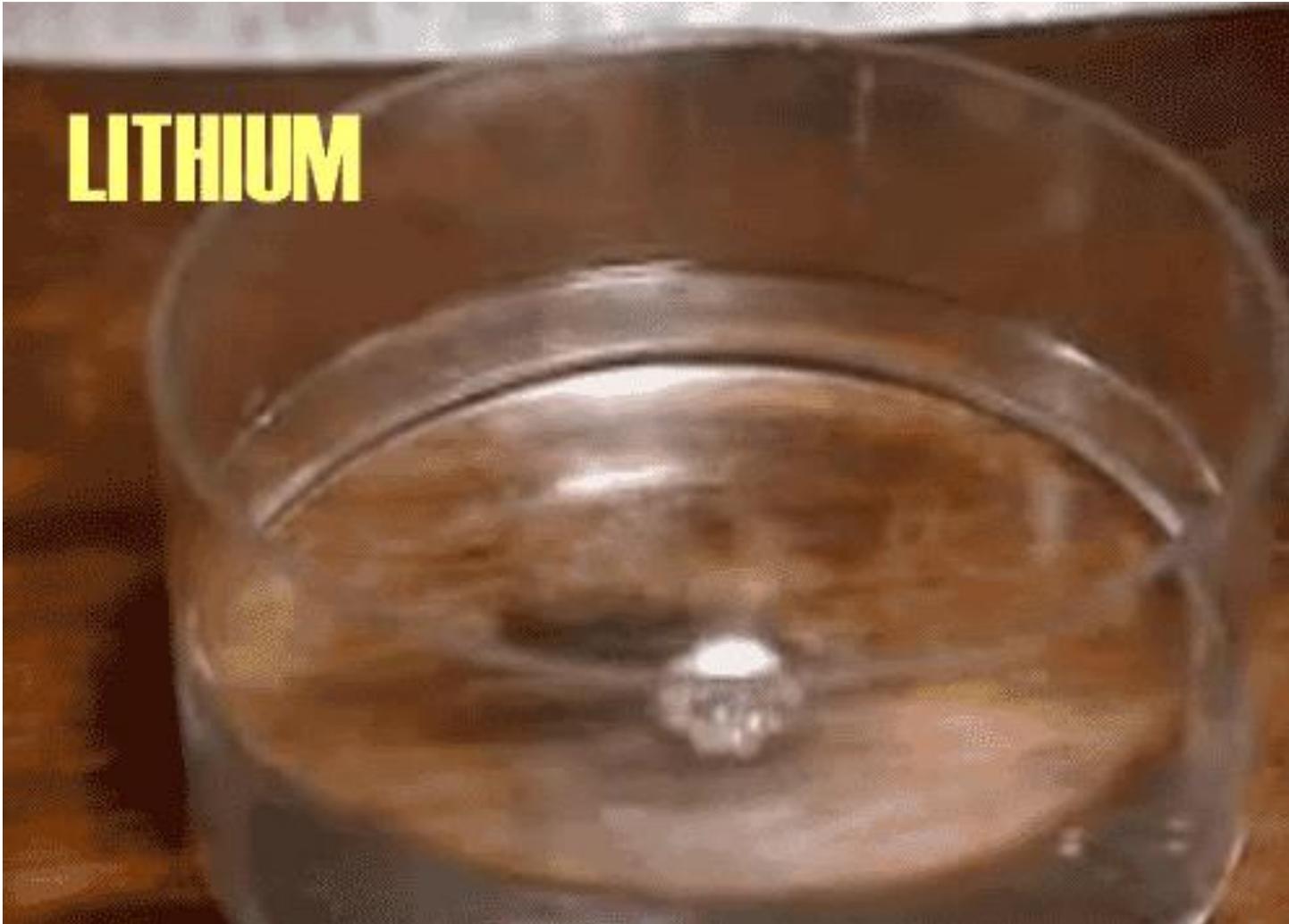
- Found in the **first column** of the periodic table
- They have **one** valence electron on their outer energy shell
- **Most reactive** metals & react violently with water
- They are **never** found as free elements in nature. They are **always bonded with another element**



Li	Lithium
Na	Sodium
K	Potassium
Rb	Rubidium
Cs	Cesium
Fr	Francium

# Group 1: Alkali Metals

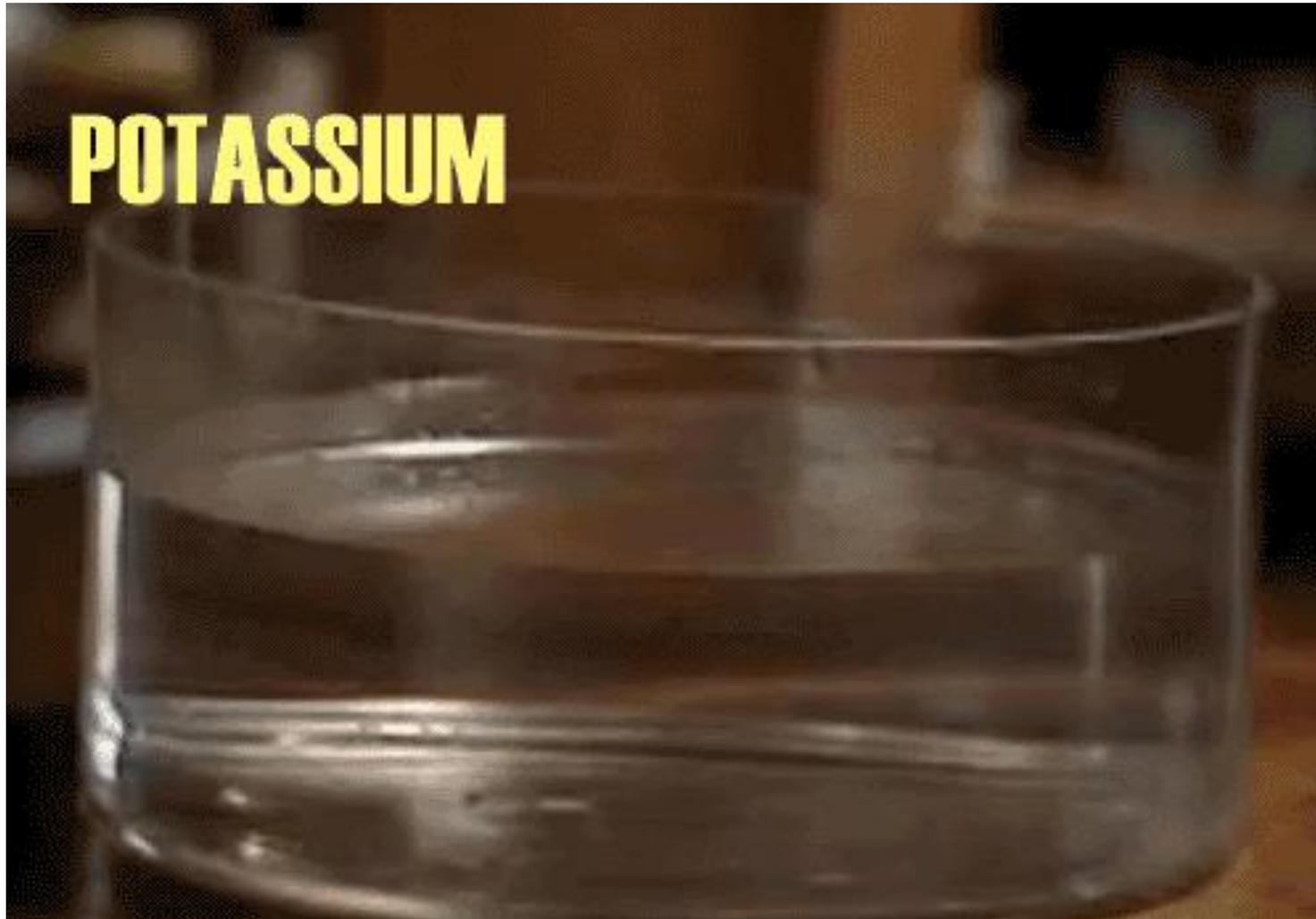
- Most reactive metals & react violently with water



Li	Lithium
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# Group 1: Alkali Metals

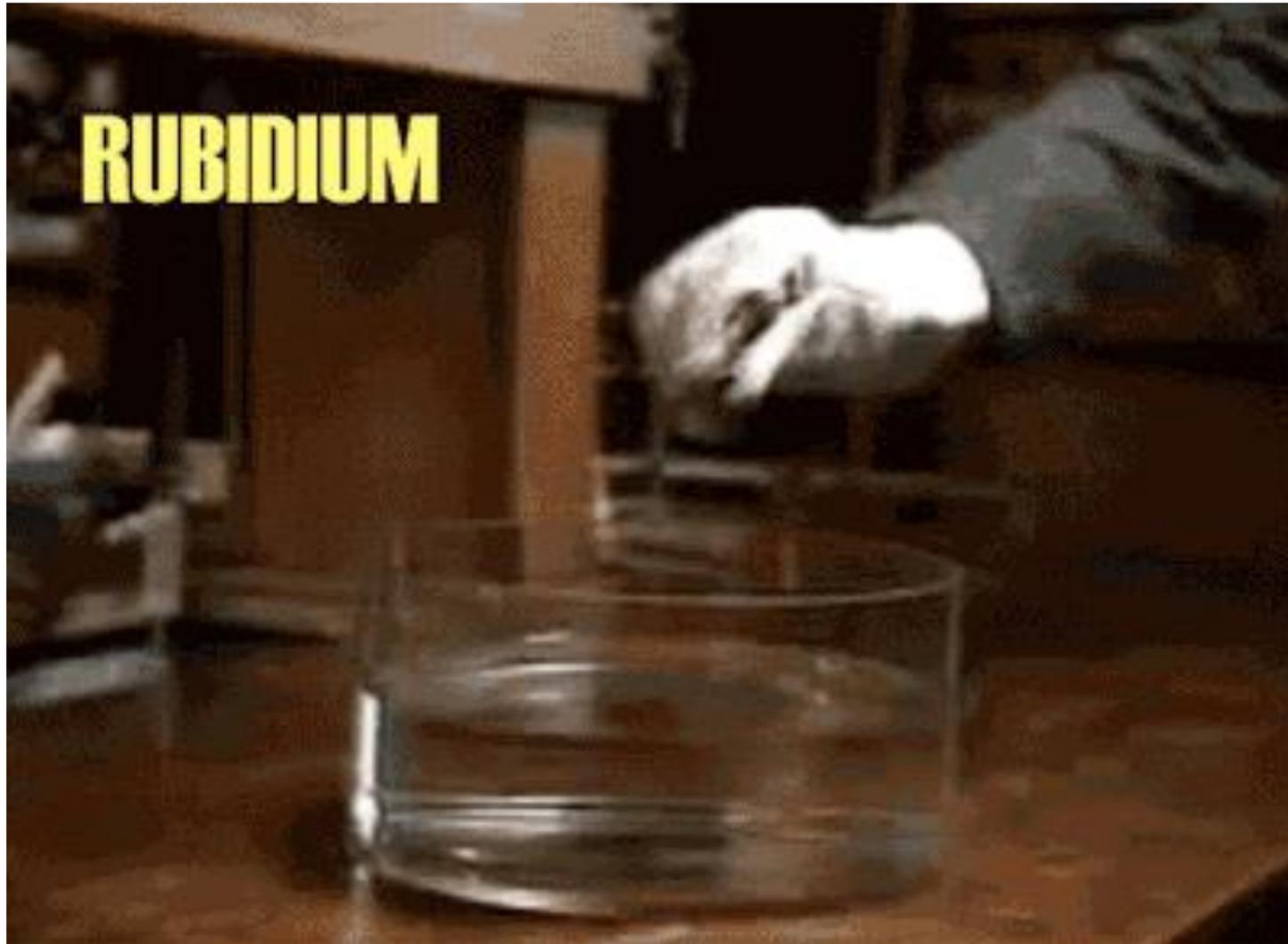
- Most reactive metals & react violently with water



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# Group 1: Alkali Metals

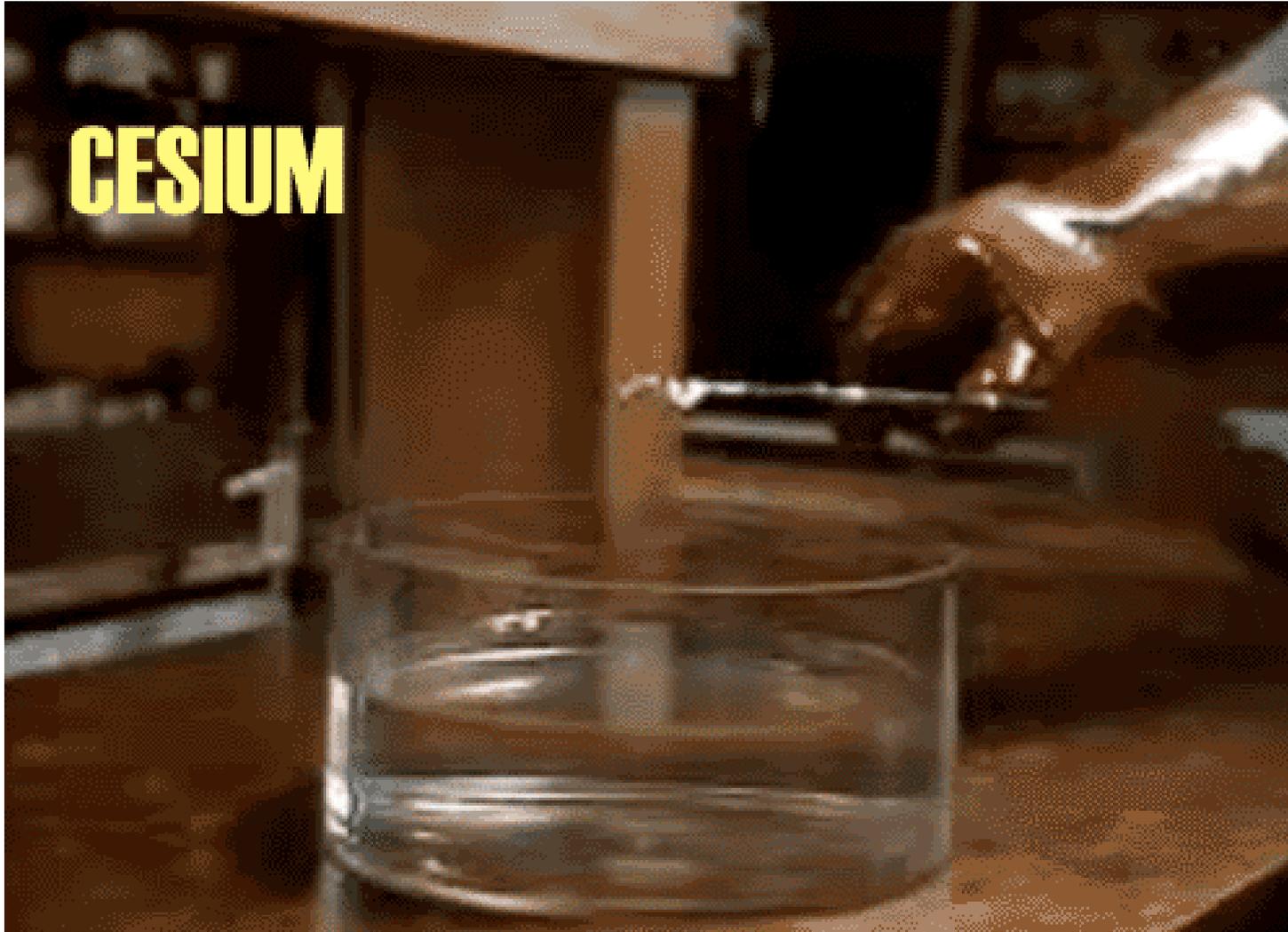
- Most reactive metals & react violently with water



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Na	Sodium
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# Group 1: Alkali Metals

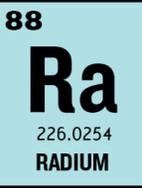
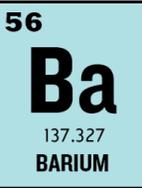
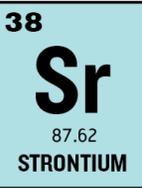
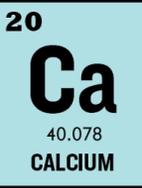
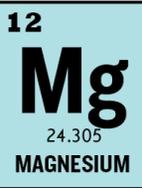
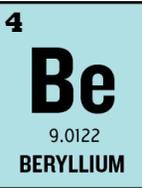
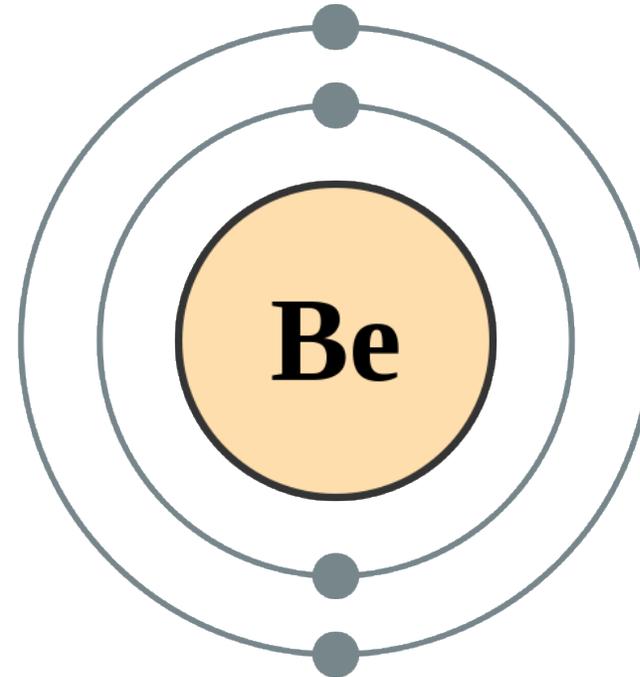
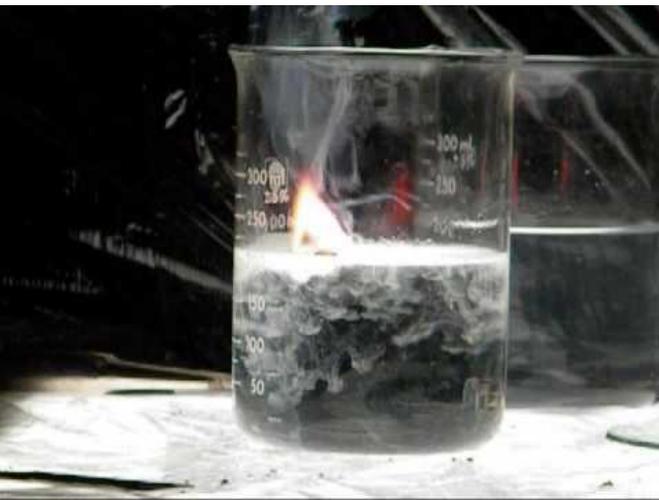
- Most reactive metals & react violently with water



Li	Lithium
Na	Sodium
K	Potassium
Rb	Rubidium
Cs	Cesium
Fr	Francium

# Group 2: Alkaline Earth Metals

- Found in the **second column** of the periodic table
- They have **two** valence electrons on their outer most energy shell
- They are also **very reactive**, but not as reactive as alkali metals
- They are **never** found free in nature. They are **always bonded with another element**



# Group 17: Halogens

- Found in the **seventeenth** column of the periodic table
- They have **seven** valence electrons in their outer most energy shell
- They are the **most active non-metals** & are **never** found free in nature

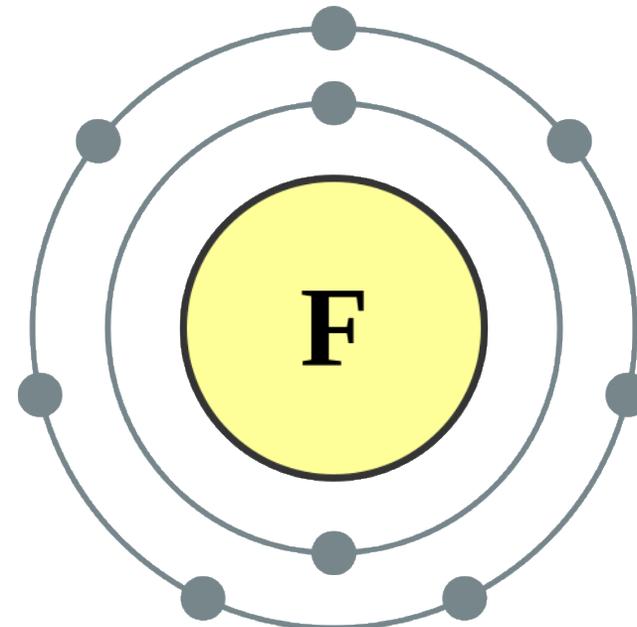
**Chlorine**



**Bromine**



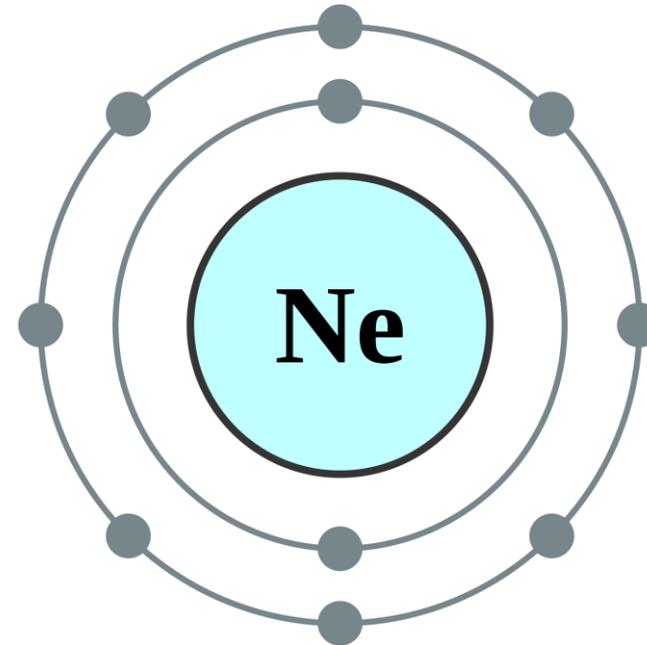
**Iodine**



9	19.0	<b>F</b> Fluorine
17	35.5	<b>Cl</b> Chlorine
35	79.9	<b>Br</b> Bromine
53	126.9	<b>I</b> Iodine
85	210	<b>At</b> Astatine

# Group 18: Noble Gases

- Found in the **eighteenth** column of the periodic table
- They have **eight** valence electrons in their outer most energy shell
- They are very **stable** and do not react with other elements (sometimes referred to as **inert**)



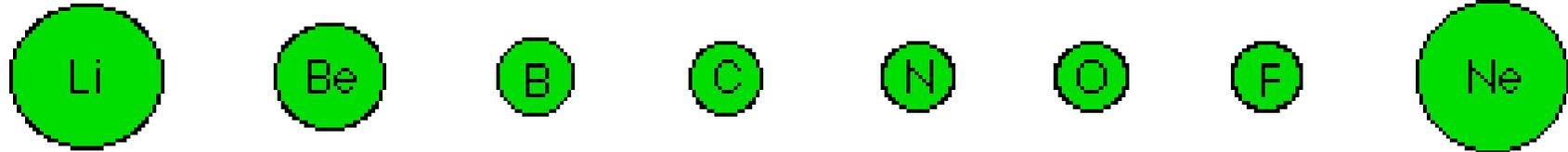
2	4
<b>He</b>	
Helium	
10	20.1
<b>Ne</b>	
Neon	
18	39.9
<b>Ar</b>	
Argon	
36	83.8
<b>Kr</b>	
Krypton	
54	131.3
<b>Xe</b>	
Xenon	
86	222
<b>Rn</b>	
Radon	

# What are Periodic Trends?

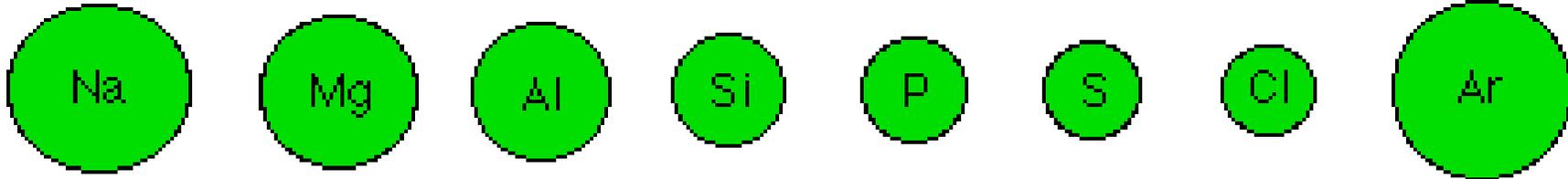
**Periodic Trend:** A general pattern that can be observed by looking at the arrangement of the elements on the periodic table.

Periodic Trends:

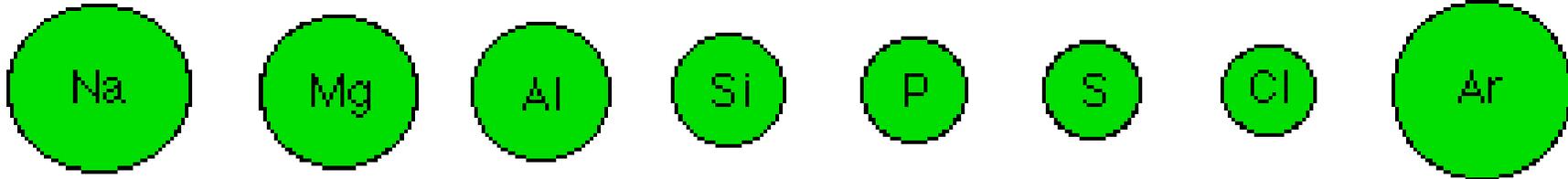
1. **Atomic Radius**



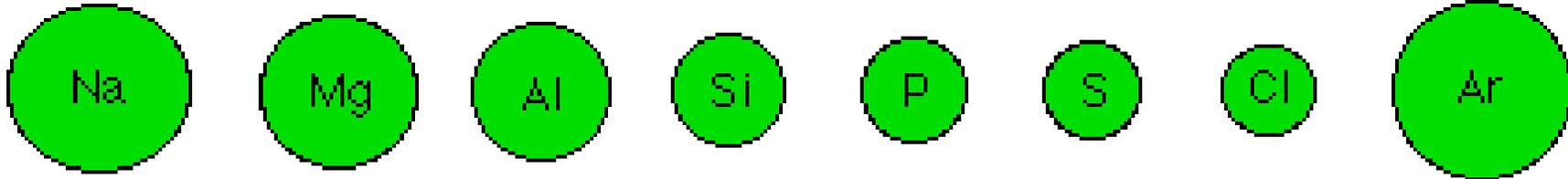
2. **Ionization Energy**



3. **Electron Affinity**



4. **Electronegativity**



5. **Electrical Conductivity**

# Atomic Radius

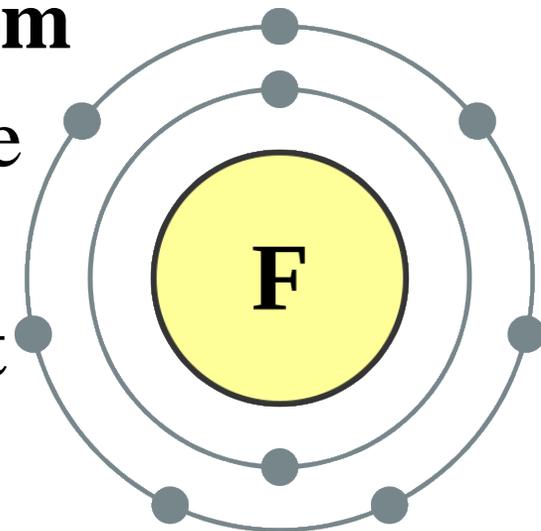
**Atomic Radius:** the size of an atom. It is the distance from the nucleus to the outermost electron.

## 1. Atomic radius increases as you go from top to bottom

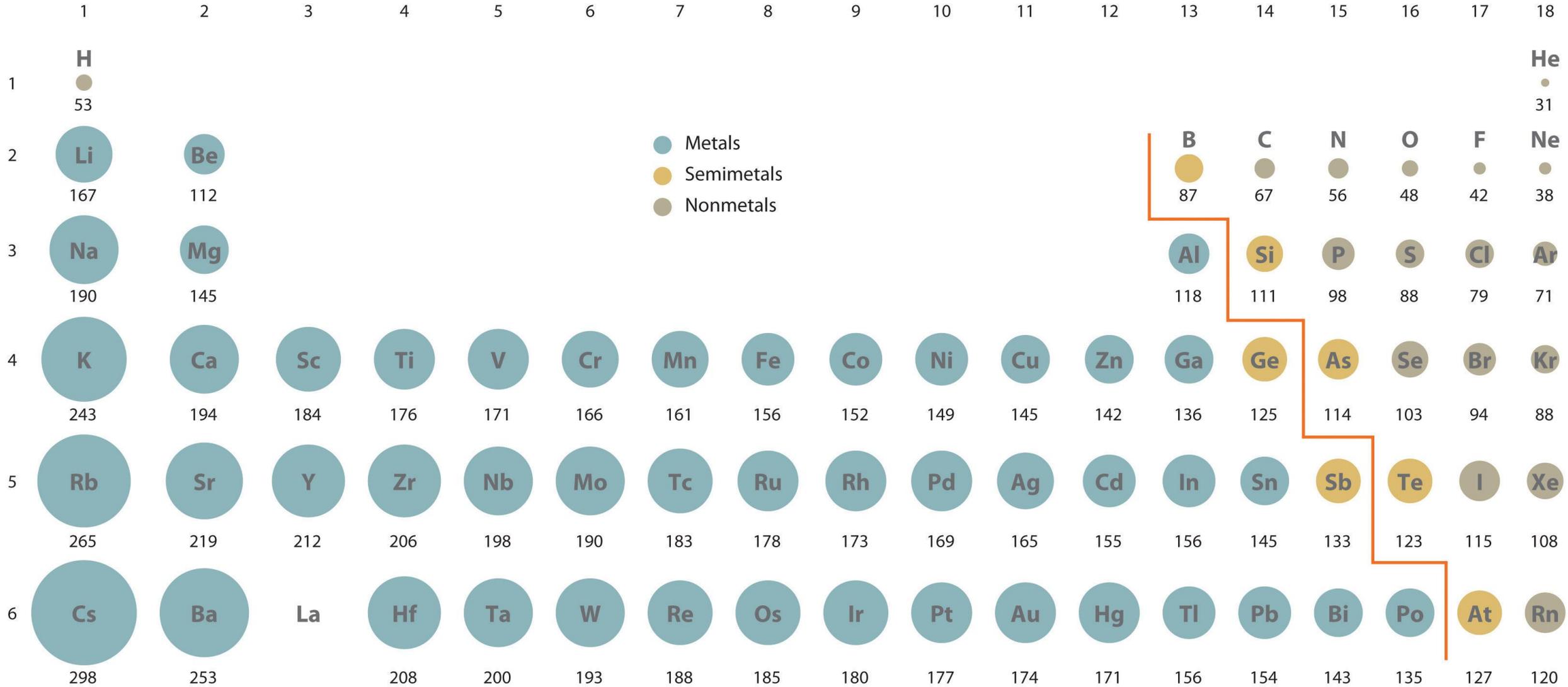
- Electrons fill energy shells that are further away from the nucleus, making the atom bigger

## 2. Atomic radius increases as you go from right to left

- Electrons fill up the same energy shell but number of proton increases, pulling the electrons closer to the nucleus



# Atomic Radius



# Atomic Radius

INCREASING ATOMIC RADIUS

INCREASING ATOMIC RADIUS

1 <b>H</b> Hydrogen 1.00794																	2 <b>He</b> Helium 4.003
3 <b>Li</b> Lithium 6.941	4 <b>Be</b> Beryllium 9.012182											5 <b>B</b> Boron 10.811	6 <b>C</b> Carbon 12.0107	7 <b>N</b> Nitrogen 14.00674	8 <b>O</b> Oxygen 15.9994	9 <b>F</b> Fluorine 18.9984632	10 <b>Ne</b> Neon 20.1797
11 <b>Na</b> Sodium 22.989770	12 <b>Mg</b> Magnesium 24.3050											13 <b>Al</b> Aluminum 26.981538	14 <b>Si</b> Silicon 28.0855	15 <b>P</b> Phosphorus 30.973761	16 <b>S</b> Sulfur 32.066	17 <b>Cl</b> Chlorine 35.4527	18 <b>Ar</b> Argon 39.948
19 <b>K</b> Potassium 39.0983	20 <b>Ca</b> Calcium 40.078	21 <b>Sc</b> Scandium 44.955910	22 <b>Ti</b> Titanium 47.867	23 <b>V</b> Vanadium 50.9415	24 <b>Cr</b> Chromium 51.9961	25 <b>Mn</b> Manganese 54.938049	26 <b>Fe</b> Iron 55.845	27 <b>Co</b> Cobalt 58.933200	28 <b>Ni</b> Nickel 58.6934	29 <b>Cu</b> Copper 63.546	30 <b>Zn</b> Zinc 65.39	31 <b>Ga</b> Gallium 69.723	32 <b>Ge</b> Germanium 72.61	33 <b>As</b> Arsenic 74.92160	34 <b>Se</b> Selenium 78.96	35 <b>Br</b> Bromine 79.904	36 <b>Kr</b> Krypton 83.80
37 <b>Rb</b> Rubidium 85.4678	38 <b>Sr</b> Strontium 87.62	39 <b>Y</b> Yttrium 88.90585	40 <b>Zr</b> Zirconium 91.224	41 <b>Nb</b> Niobium 92.90638	42 <b>Mo</b> Molybdenum 95.94	43 <b>Tc</b> Technetium (98)	44 <b>Ru</b> Ruthenium 101.07	45 <b>Rh</b> Rhodium 102.90550	46 <b>Pd</b> Palladium 106.42	47 <b>Ag</b> Silver 107.8682	48 <b>Cd</b> Cadmium 112.411	49 <b>In</b> Indium 114.818	50 <b>Sn</b> Tin 118.710	51 <b>Sb</b> Antimony 121.760	52 <b>Te</b> Tellurium 127.60	53 <b>I</b> Iodine 126.90447	54 <b>Xe</b> Xenon 131.29
55 <b>Cs</b> Cesium 132.90545	56 <b>Ba</b> Barium 137.327	57 <b>La</b> Lanthanum 138.9055	72 <b>Hf</b> Hafnium 178.49	73 <b>Ta</b> Tantalum 180.9479	74 <b>W</b> Tungsten 183.84	75 <b>Re</b> Rhenium 186.207	76 <b>Os</b> Osmium 190.23	77 <b>Ir</b> Iridium 192.217	78 <b>Pt</b> Platinum 195.078	79 <b>Au</b> Gold 196.96655	80 <b>Hg</b> Mercury 200.59	81 <b>Tl</b> Thallium 204.3833	82 <b>Pb</b> Lead 207.2	83 <b>Bi</b> Bismuth 208.98038	84 <b>Po</b> Polonium (209)	85 <b>At</b> Astatine (210)	86 <b>Rn</b> Radon (222)
87 <b>Fr</b> Francium (223)	88 <b>Ra</b> Radium (226)	89 <b>Ac</b> Actinium (227)	104 <b>Rf</b> Rutherfordium (261)	105 <b>Db</b> Dubnium (262)	106 <b>Sg</b> Seaborgium (263)	107 <b>Bh</b> Bohrium (262)	108 <b>Hs</b> Hassium (265)	109 <b>Mt</b> Meitnerium (266)	110 (268)	111 (272)	112 (277)	113	114				

# **Think, Pair & Share**

**How do you think the size of the Cation will be different from the same element's neutral atom?**

**How do you think the size of the anion will be different from the same element's neutral atom?**

A large, empty rectangular box with a thick black border, intended for students to write their answers to the questions above.

# Think, Pair & Share

How do you think the size of the Cation will be different from the same element's neutral atom?

How do you think the size of the anion will be different from the same element's neutral atom?

**Cation:** Will lose electrons and therefore will have a smaller radius

# Think, Pair & Share

How do you think the size of the Cation will be different from the same element's neutral atom?

How do you think the size of the anion will be different from the same element's neutral atom?

**Cation:** Will lose electrons and therefore will have a smaller radius

**Anion:** Will gain electrons and therefore will have a larger radius

# **Ionization Energy**

**Ionization Energy:** The amount of energy needed to remove the outermost electron

## **1. Ionization energy increases as you go from bottom to top**

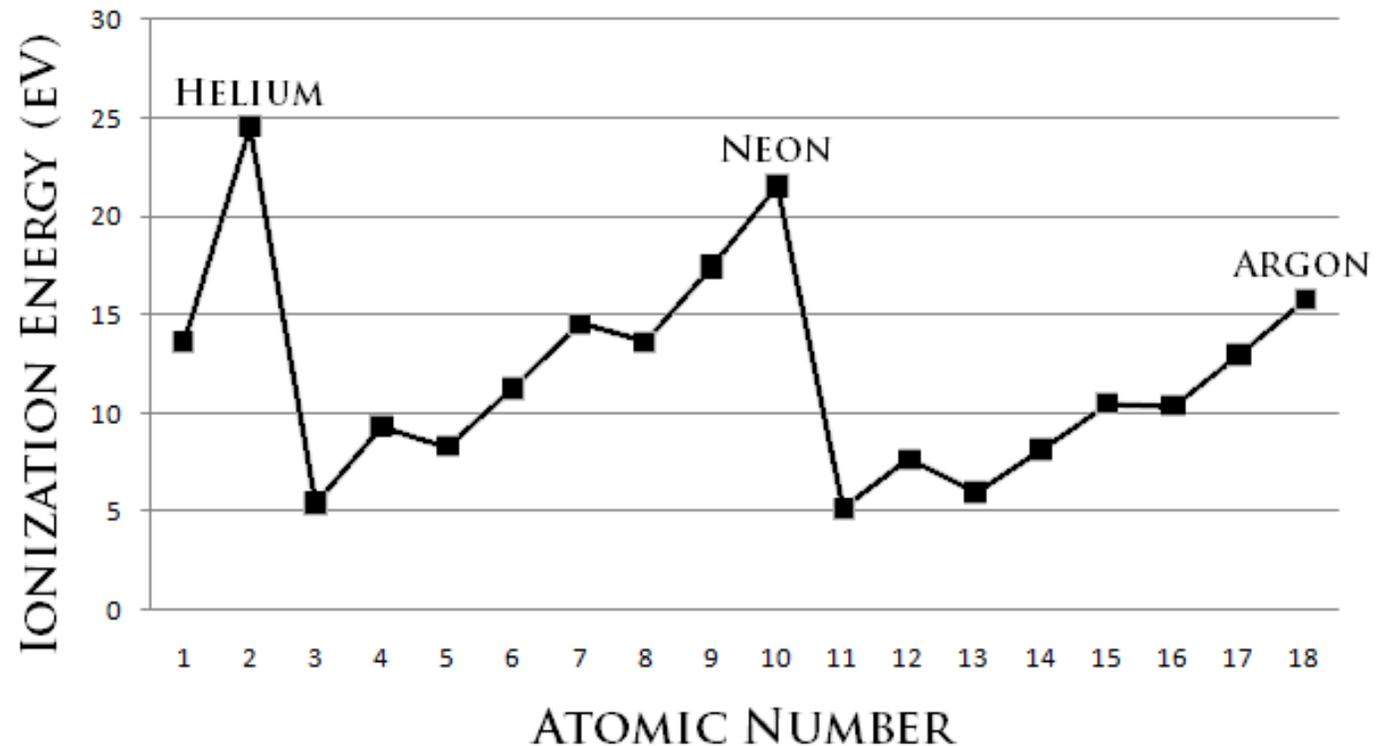
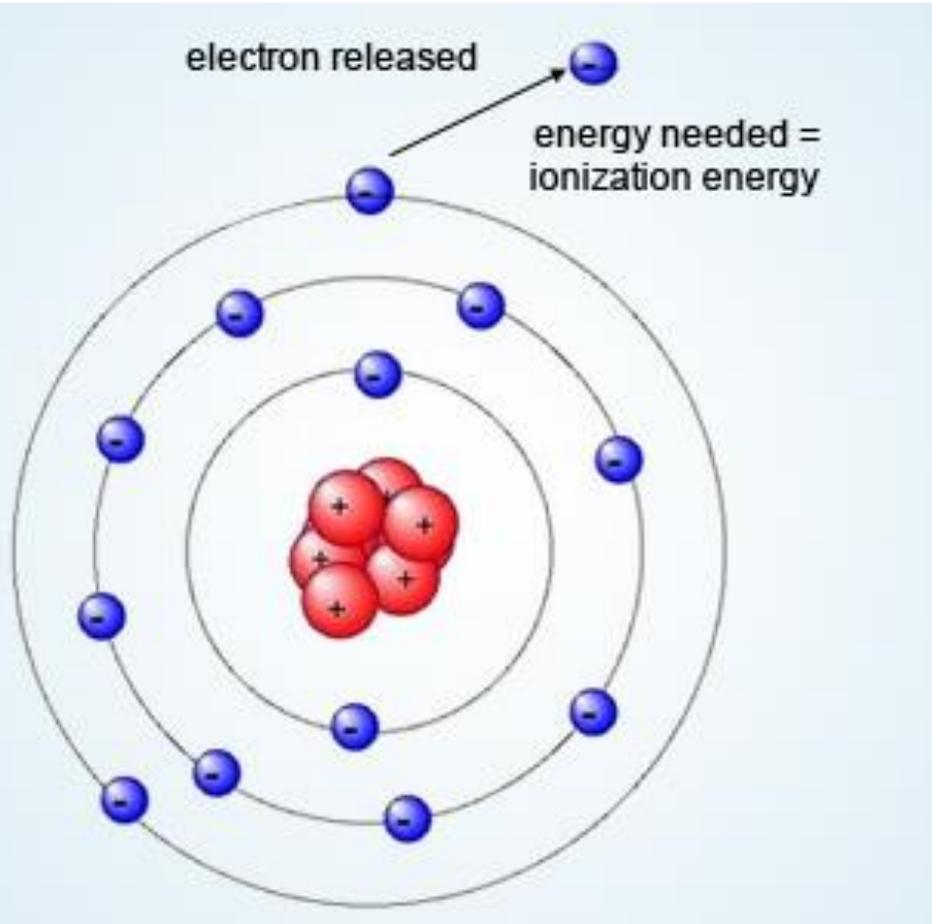
- As you go from **bottom to top**, the atomic size **decreases** so the electrons are closer to the nucleus. Therefore, it is harder to remove

## **2. Ionization energy increase as you go from left to right**

- As you go from left to right, the atomic size decreases so the electrons are closer to the nucleus. Therefore, it is harder to remove

# Ionization Energy

- The energy needed to remove the outermost electron



# Ionization Energy

Increasing Ionization Energy

High Energy

	IA												IIIA	IVA	VA	VIA	VIIA	VIIIA
1	H	IIA																He
2	Li	Be											B	C	N	O	F	Ne
3	Na	Mg	IIIB	IVB	VB	VIB	VIIIB	VIII B			IB	IIB	Al	Si	P	S	Cl	Ar
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
6	Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
7	Fr	Rd	Ac															

Low Energy

Increasing Ionization Energy

# Think, Pair & Share

How do you think ionization energy of a metal affects how reactive it is?

A large, empty rectangular box with a thick black border, intended for students to write their answers to the question above.

# Think, Pair & Share

How do you think ionization energy of a metal affects how reactive it is?

The **lower** the ionization energy, the more reactive the element will be. Think about Alkali metals.

# **Electron Affinity**

**Electron Affinity:** The likelihood of an atom to gain an electron

## **1. Electron affinity increases as you go from bottom to top**

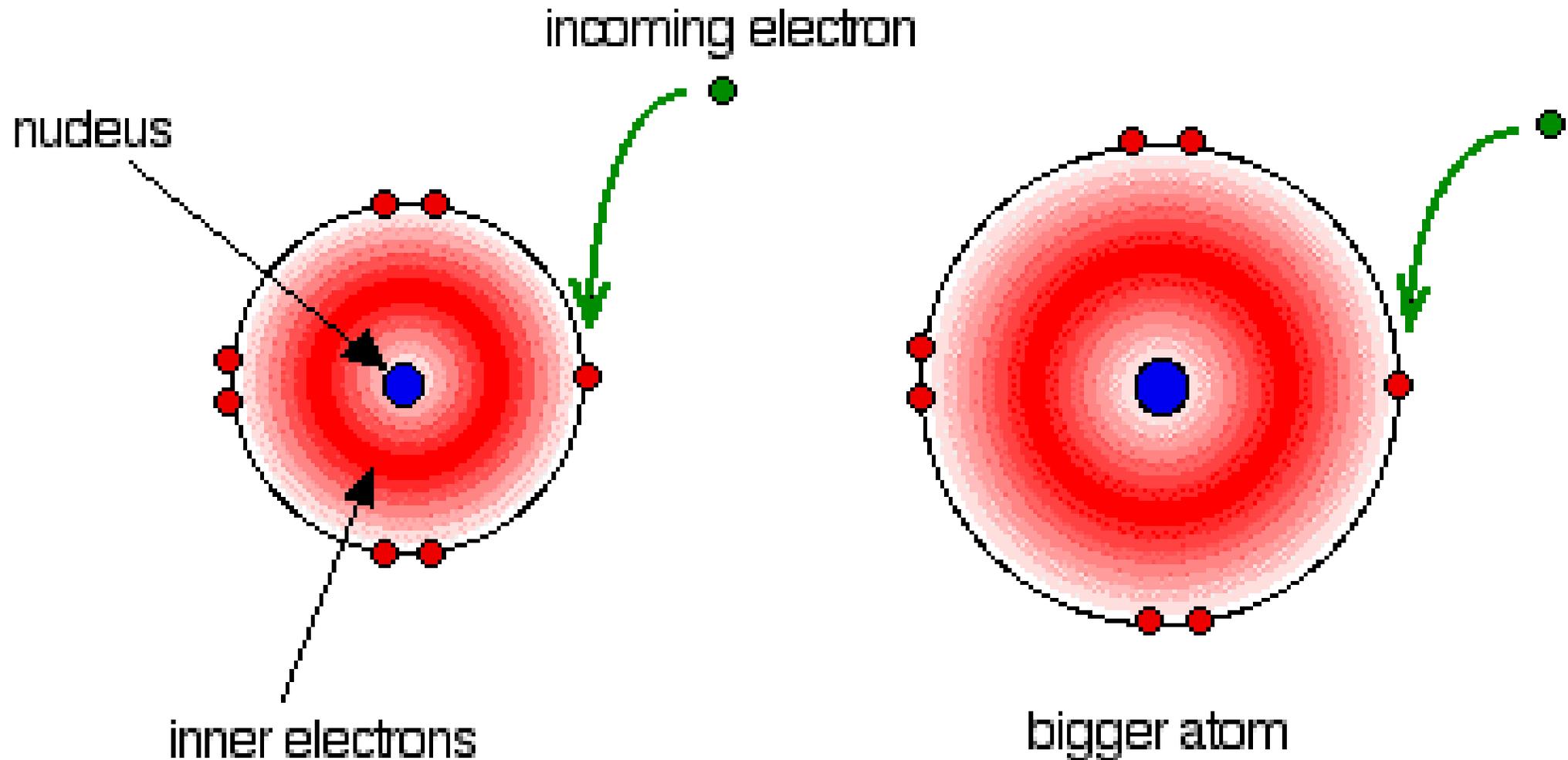
- As you go from bottom to top, the atomic size decreases so the electrons are closer to the nucleus. Therefore, it is easier to add an electron

## **2. Electron affinity increases as you go from left to right**

- As you go from left to right, the atomic size decreases so the electrons are closer to the nucleus. Therefore, it is easier to add an electron

# Electron Affinity

The likelihood of gaining an electron



# Electron Affinity

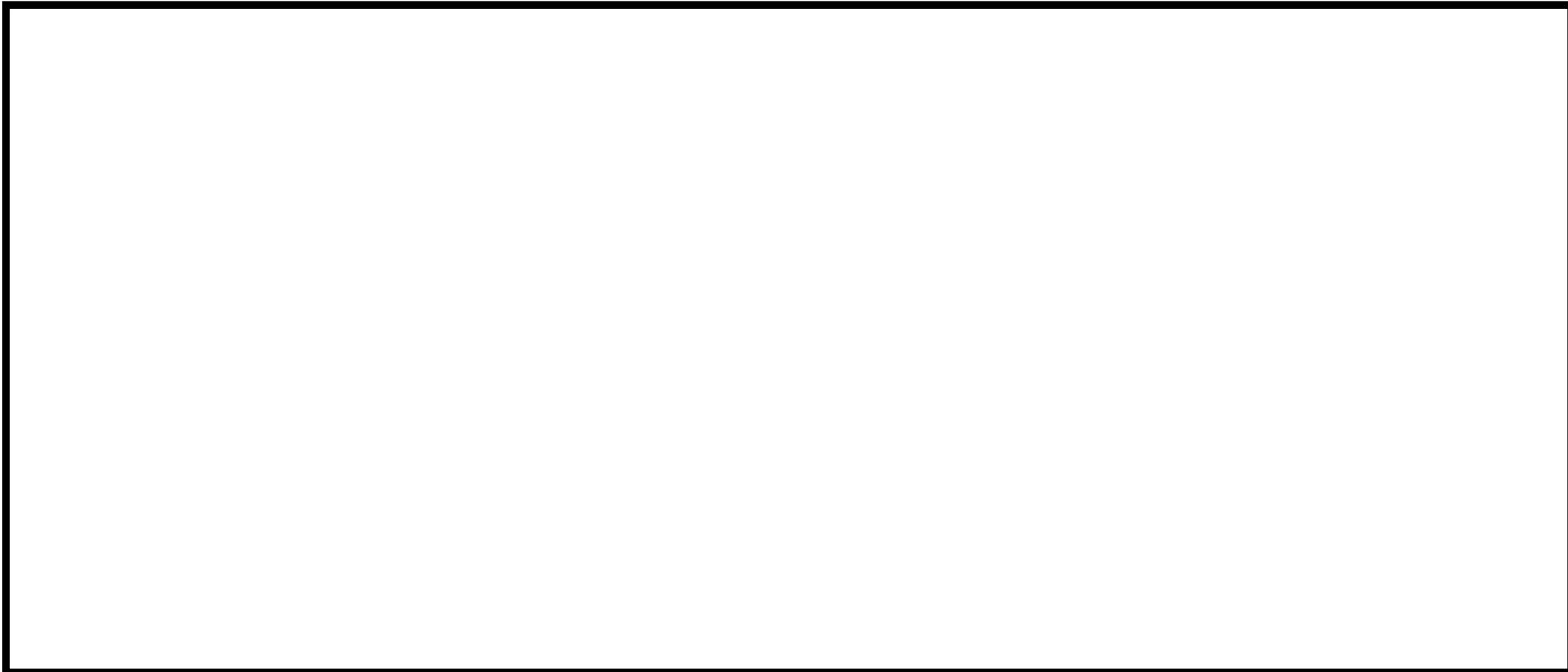
## INCREASING ELECTRON AFFINITY

1 <b>H</b> Hydrogen 1.00794																	2 <b>He</b> Helium 4.003
3 <b>Li</b> Lithium 6.941	4 <b>Be</b> Beryllium 9.012182											5 <b>B</b> Boron 10.811	6 <b>C</b> Carbon 12.0107	7 <b>N</b> Nitrogen 14.00674	8 <b>O</b> Oxygen 15.9994	9 <b>F</b> Fluorine 18.9984032	10 <b>Ne</b> Neon 20.1797
11 <b>Na</b> Sodium 22.989770	12 <b>Mg</b> Magnesium 24.3050											13 <b>Al</b> Aluminum 26.981538	14 <b>Si</b> Silicon 28.0855	15 <b>P</b> Phosphorus 30.973761	16 <b>S</b> Sulfur 32.066	17 <b>Cl</b> Chlorine 35.4527	18 <b>Ar</b> Argon 39.948
19 <b>K</b> Potassium 39.0983	20 <b>Ca</b> Calcium 40.078	21 <b>Sc</b> Scandium 44.955910	22 <b>Ti</b> Titanium 47.867	23 <b>V</b> Vanadium 50.9415	24 <b>Cr</b> Chromium 51.9961	25 <b>Mn</b> Manganese 54.938049	26 <b>Fe</b> Iron 55.845	27 <b>Co</b> Cobalt 58.933200	28 <b>Ni</b> Nickel 58.6934	29 <b>Cu</b> Copper 63.546	30 <b>Zn</b> Zinc 65.39	31 <b>Ga</b> Gallium 69.723	32 <b>Ge</b> Germanium 72.61	33 <b>As</b> Arsenic 74.92160	34 <b>Se</b> Selenium 78.96	35 <b>Br</b> Bromine 79.904	36 <b>Kr</b> Krypton 83.80
37 <b>Rb</b> Rubidium 85.4678	38 <b>Sr</b> Strontium 87.62	39 <b>Y</b> Yttrium 88.90585	40 <b>Zr</b> Zirconium 91.224	41 <b>Nb</b> Niobium 92.90638	42 <b>Mo</b> Molybdenum 95.94	43 <b>Tc</b> Technetium (98)	44 <b>Ru</b> Ruthenium 101.07	45 <b>Rh</b> Rhodium 102.90550	46 <b>Pd</b> Palladium 106.42	47 <b>Ag</b> Silver 107.8682	48 <b>Cd</b> Cadmium 112.411	49 <b>In</b> Indium 114.818	50 <b>Sn</b> Tin 118.710	51 <b>Sb</b> Antimony 121.760	52 <b>Te</b> Tellurium 127.60	53 <b>I</b> Iodine 126.90447	54 <b>Xe</b> Xenon 131.29
55 <b>Cs</b> Cesium 132.90545	56 <b>Ba</b> Barium 137.327	57 <b>La</b> Lanthanum 138.9055	72 <b>Hf</b> Hafnium 178.49	73 <b>Ta</b> Tantalum 180.9479	74 <b>W</b> Tungsten 183.84	75 <b>Re</b> Rhenium 186.207	76 <b>Os</b> Osmium 190.23	77 <b>Ir</b> Iridium 192.217	78 <b>Pt</b> Platinum 195.078	79 <b>Au</b> Gold 196.96655	80 <b>Hg</b> Mercury 200.59	81 <b>Tl</b> Thallium 204.3833	82 <b>Pb</b> Lead 207.2	83 <b>Bi</b> Bismuth 208.98038	84 <b>Po</b> Polonium (209)	85 <b>At</b> Astatine (210)	86 <b>Rn</b> Radon (222)
87 <b>Fr</b> Francium (223)	88 <b>Ra</b> Radium (226)	89 <b>Ac</b> Actinium (227)	104 <b>Rf</b> Rutherfordium (261)	105 <b>Db</b> Dubnium (262)	106 <b>Sg</b> Seaborgium (263)	107 <b>Bh</b> Bohrium (262)	108 <b>Hs</b> Hassium (265)	109 <b>Mt</b> Meitnerium (266)	110 (269)	111 (272)	112 (277)	113	114				

INCREASING ELECTRON AFFINITY

# **Think, Pair & Share**

**How do you think the electron affinity of a non-metal affects how reactive it is?**

A large, empty rectangular box with a thick black border, intended for students to write their responses to the question above.

# Think, Pair & Share

How do you think the electron affinity of a non-metal affects how reactive it is?

- Reactivity increases from bottom to top because the atomic radius decreases from bottom to top and so the valence shell is closer to the nucleus.
- This makes smaller atoms have higher electron affinity, and increasing their reactivity.

# Recall: Electronegativity

- **Electronegativity**: the tendency of the atom to attract electrons from a neighboring atom
- Electronegativity **increases** as you go from bottom to top
- Electronegativity **increases** as you go left to right

hydrogen 1 <b>H</b> 1.0079																	helium 2 <b>He</b> 4.0026						
lithium 3 <b>Li</b> 6.941	beryllium 4 <b>Be</b> 9.0122																	boron 5 <b>B</b> 10.811	carbon 6 <b>C</b> 12.011	nitrogen 7 <b>N</b> 14.007	oxygen 8 <b>O</b> 15.999	fluorine 9 <b>F</b> 18.998	neon 10 <b>Ne</b> 20.180
sodium 11 <b>Na</b> 22.990	magnesium 12 <b>Mg</b> 24.305																	aluminum 13 <b>Al</b> 26.982	silicon 14 <b>Si</b> 28.086	phosphorus 15 <b>P</b> 30.974	sulfur 16 <b>S</b> 32.065	chlorine 17 <b>Cl</b> 35.453	argon 18 <b>Ar</b> 39.948
potassium 19 <b>K</b> 39.098	calcium 20 <b>Ca</b> 40.078	scandium 21 <b>Sc</b> 44.956	titanium 22 <b>Ti</b> 47.867	vanadium 23 <b>V</b> 50.942	chromium 24 <b>Cr</b> 51.996	manganese 25 <b>Mn</b> 54.938	iron 26 <b>Fe</b> 55.845	cobalt 27 <b>Co</b> 58.933	nickel 28 <b>Ni</b> 58.693	copper 29 <b>Cu</b> 63.546	zinc 30 <b>Zn</b> 65.39	gallium 31 <b>Ga</b> 69.723	germanium 32 <b>Ge</b> 72.61	arsenic 33 <b>As</b> 74.922	selenium 34 <b>Se</b> 78.96	bromine 35 <b>Br</b> 79.904	krypton 36 <b>Kr</b> 83.80						
rubidium 37 <b>Rb</b> 85.468	strontium 38 <b>Sr</b> 87.62	yttrium 39 <b>Y</b> 88.906	zirconium 40 <b>Zr</b> 91.224	niobium 41 <b>Nb</b> 92.906	molybdenum 42 <b>Mo</b> 95.94	technetium 43 <b>Tc</b> [98]	ruthenium 44 <b>Ru</b> 101.07	rhodium 45 <b>Rh</b> 102.91	palladium 46 <b>Pd</b> 106.42	silver 47 <b>Ag</b> 107.87	cadmium 48 <b>Cd</b> 112.41	indium 49 <b>In</b> 114.82	tin 50 <b>Sn</b> 118.71	antimony 51 <b>Sb</b> 121.76	tellurium 52 <b>Te</b> 127.60	iodine 53 <b>I</b> 126.90	xenon 54 <b>Xe</b> 131.29						
caesium 55 <b>Cs</b> 132.91	barium 56 <b>Ba</b> 137.33	57-70 *	lutetium 71 <b>Lu</b> 174.97	hafnium 72 <b>Hf</b> 178.49	tantalum 73 <b>Ta</b> 180.95	tungsten 74 <b>W</b> 183.84	rhenium 75 <b>Re</b> 186.21	osmium 76 <b>Os</b> 190.23	iridium 77 <b>Ir</b> 192.22	platinum 78 <b>Pt</b> 195.08	gold 79 <b>Au</b> 196.97	mercury 80 <b>Hg</b> 200.59	thallium 81 <b>Tl</b> 204.38	lead 82 <b>Pb</b> 207.2	bismuth 83 <b>Bi</b> 208.98	polonium 84 <b>Po</b> [209]	astatine 85 <b>At</b> [210]	radon 86 <b>Rn</b> [222]					
francium 87 <b>Fr</b> [223]	radium 88 <b>Ra</b> [226]	89-102 * *	lawrencium 103 <b>Lr</b> [262]	rutherfordium 104 <b>Rf</b> [261]	dubnium 105 <b>Db</b> [262]	seaborgium 106 <b>Sg</b> [263]	bohrium 107 <b>Bh</b> [264]	hassium 108 <b>Hs</b> [269]	meitnerium 109 <b>Mt</b> [268]	unnilium 110 <b>Uun</b> [271]	ununium 111 <b>Uuu</b> [272]	unbibium 112 <b>Uub</b> [277]	ununquadium 114 <b>Uuq</b> [289]										

\* Lanthanide series

lanthanum 57 <b>La</b> 138.91	cerium 58 <b>Ce</b> 140.12	praseodymium 59 <b>Pr</b> 140.91	neodymium 60 <b>Nd</b> 144.24	promethium 61 <b>Pm</b> [145]	samarium 62 <b>Sm</b> 150.36	europium 63 <b>Eu</b> 151.96	gadolinium 64 <b>Gd</b> 157.25	terbium 65 <b>Tb</b> 158.93	dysprosium 66 <b>Dy</b> 162.50	holmium 67 <b>Ho</b> 164.93	erbium 68 <b>Er</b> 167.26	thulium 69 <b>Tm</b> 168.93	ytterbium 70 <b>Yb</b> 173.04
actinium 89 <b>Ac</b> [227]	thorium 90 <b>Th</b> 232.04	protactinium 91 <b>Pa</b> 231.04	uranium 92 <b>U</b> 238.03	neptunium 93 <b>Np</b> [237]	plutonium 94 <b>Pu</b> [244]	americium 95 <b>Am</b> [243]	curium 96 <b>Cm</b> [247]	berkelium 97 <b>Bk</b> [247]	californium 98 <b>Cf</b> [251]	einsteinium 99 <b>Es</b> [252]	fermium 100 <b>Fm</b> [257]	mendelevium 101 <b>Md</b> [258]	nobelium 102 <b>No</b> [259]

\*\* Actinide series

# Think, Pair & Share

How is electronegativity related to ionization energy?

A large, empty rectangular box with a thick black border, intended for students to write their response to the question above.

# Think, Pair & Share

How is electronegativity related to ionization energy?

If an atom has a high electronegativity, then that atom is highly attracted to its own valence electrons.

As a result, it will take a lot of energy to remove its valence electrons and therefore will have a high ionization energy.

# Recall: Electronegativity

**INCREASING** 

hydrogen 1 <b>H</b> 1.0079																	helium 2 <b>He</b> 4.0026						
lithium 3 <b>Li</b> 6.941	beryllium 4 <b>Be</b> 9.0122																	boron 5 <b>B</b> 10.811	carbon 6 <b>C</b> 12.011	nitrogen 7 <b>N</b> 14.007	oxygen 8 <b>O</b> 15.999	fluorine 9 <b>F</b> 18.998	neon 10 <b>Ne</b> 20.180
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potassium 19 <b>K</b> 39.098	calcium 20 <b>Ca</b> 40.078	scandium 21 <b>Sc</b> 44.956	titanium 22 <b>Ti</b> 47.867	vanadium 23 <b>V</b> 50.942	chromium 24 <b>Cr</b> 51.996	manganese 25 <b>Mn</b> 54.938	iron 26 <b>Fe</b> 55.845	cobalt 27 <b>Co</b> 58.933	nickel 28 <b>Ni</b> 58.693	copper 29 <b>Cu</b> 63.546	zinc 30 <b>Zn</b> 65.39	gallium 31 <b>Ga</b> 69.723	germanium 32 <b>Ge</b> 72.61	arsenic 33 <b>As</b> 74.922	selenium 34 <b>Se</b> 78.96	bromine 35 <b>Br</b> 79.904	krypton 36 <b>Kr</b> 83.80						
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\* Lanthanide series

\*\* Actinide series

lanthanum 57 <b>La</b> 138.91	cerium 58 <b>Ce</b> 140.12	praseodymium 59 <b>Pr</b> 140.91	neodymium 60 <b>Nd</b> 144.24	promethium 61 <b>Pm</b> [145]	samarium 62 <b>Sm</b> 150.36	europium 63 <b>Eu</b> 151.96	gadolinium 64 <b>Gd</b> 157.25	terbium 65 <b>Tb</b> 158.93	dysprosium 66 <b>Dy</b> 162.50	holmium 67 <b>Ho</b> 164.93	erbium 68 <b>Er</b> 167.26	thulium 69 <b>Tm</b> 168.93	ytterbium 70 <b>Yb</b> 173.04
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# **Electrical Conductivity**

**Electrical Conductivity:** The ability of a material to **conduct** (transport) electric current

## **1. Electrical conductivity increases as you go from top to bottom**

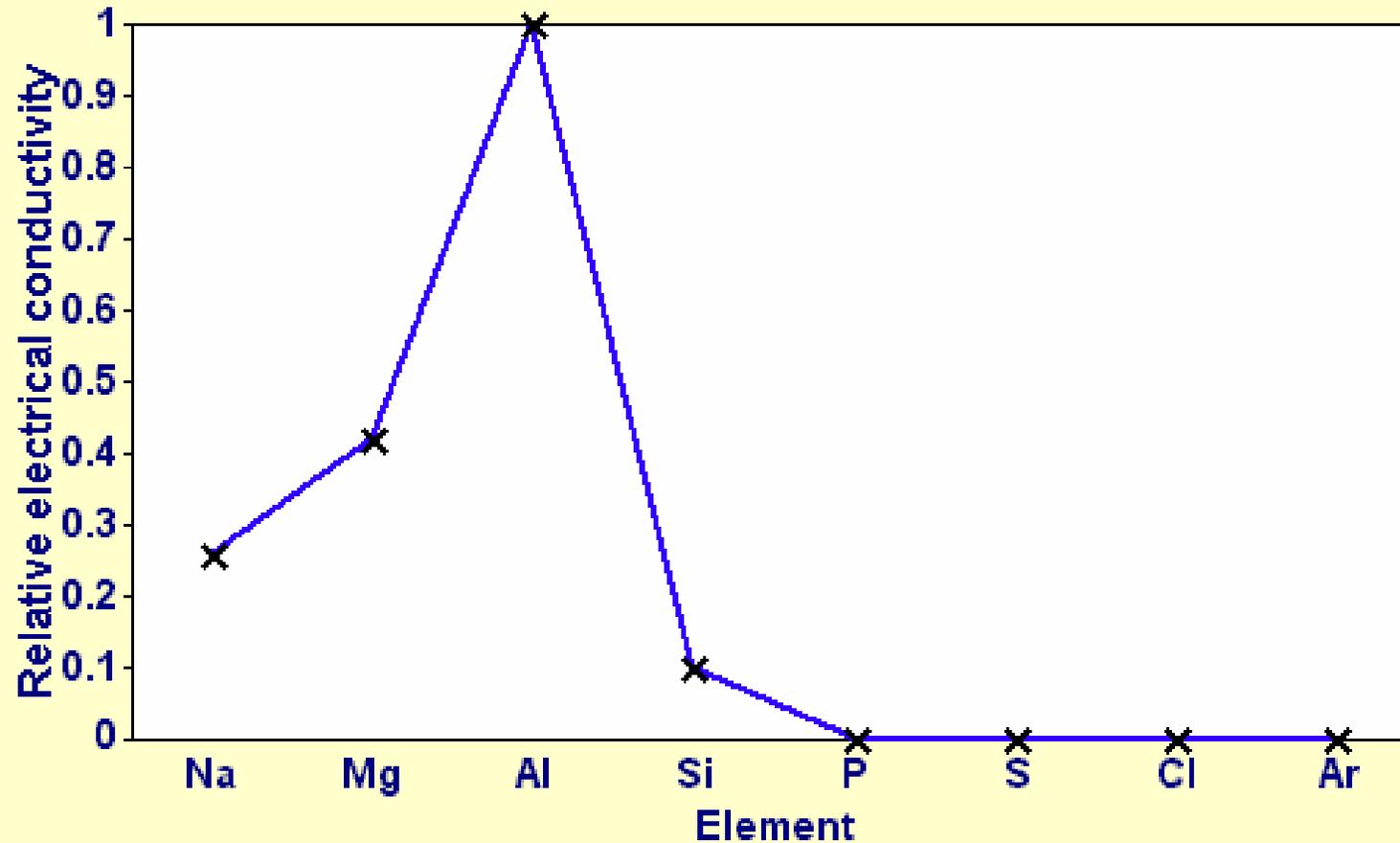
- As you go from top to bottom, the atomic size increases & the outermost electrons are further from the nucleus. Therefore, they are more easily able to move around & transport electric current

## **2. Electrical conductivity increases as you go from right to left**

- As you go from right to left, the atomic size increases & the outermost electrons are further from the nucleus. Therefore, they are more easily able to move around & transport electric current

# Electrical Conductivity

Trend in electrical conductivity in Period 3



# Periodic Trends Summary

